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Hint: add your answer to question 1 and four times your answer to question 2.  $4 + 4(7) = 4 + 28 = 32$  4. Verify that there are 32 electrons pictured in the drawing of  $\text{CCl}_4$ . Count each dot in the picture above: there are 24. There are also 8 more electrons represented by the 4 lines (each line represents two electrons). 5.

ChemQuest 23 : Drawing Covalent Compounds

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Hint: add your answer to question 1 and four times your answer to question 2.  $n\$ t ?(q /= BZ e^*eeAronS$  Verify that there are 32 electrons pictured in the drawing of  $\text{CCl}_q$ . 3es Find the sum of all the valence electrons for  $\text{NF}.$  (Add how many valence electrons one nitrogen atom has with the valence electrons for three fluorine atoms.) H: ?crl 2 Tte /je.

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Chemquest 23 Lewis Structures Answers

Answer: + that S has a negative 2 charge. There are two "X" ions to balance out the negative 2 charge of S. Step 2: Criss cross to balance charges. Step 1: Write each ion and each charge. Positive first! Here's what you get: Add parentheses balance charges. Step 2. Criss cross to

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